

ATMOSPHERIC SALT CONFIRMED

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The unusual presence of an alkaline salt form in the atmosphere is in the process of being established and confirmed through electrolysis testing. It has been deduced and established from earlier work and research that a case for testing the atmosphere, water and soil for alkaline salts exists based upon meteorological, chemical and physical principles. Please refer to [A Case for Testing, Eight Conditions, Drastic pH Changes, pH Test Alert, 20 Times](#) and [pH Test Results](#) for additional reference material on this topic.

It is established that testing shows an approximate twenty fold increase with respect to baseline values in the level of hydroxide (OH⁻ ions) of rainwater during the previous year. These tests are statistically significant at the 99.99%+ level. This increase is directly associated with extensive and sustained aircraft aerosol operations across the country since the early part of 1999. This equivalently means that the pH of the atmosphere has recently been significantly and rapidly altered

as an apparent consequence of the aerosol operations.

The presence of an anticipated alkaline salt form, or electrolyte, has now been confirmed through electrolysis testing. The specific salt form is not yet identified, although the number of candidates is relatively few in number.

Two simultaneous conditions exist which bear out the conclusions being reached on this page. First, the dramatic increase in rainfall pH indicates that an unusual presence of hydroxide ions, (OH⁻) now commonly exists in reported rain samples across the country. Secondly, electrical current applied to these same samples results in a chemical reaction that can occur only in the presence of an electrolyte, i.e., a salt form. The specific reaction which has been recorded involves evolution of hydrogen gas at the cathode, and the dissolution of copper at the anode when copper electrodes are used. The magnitude of the reactions recorded, in comparison to control tests, indicates the presence of a metallic alkaline salt within the rainfall that has been tested.

The candidate list of salt forms is expected to consider those elements within Group I and II of the periodic table, with a special emphasis applied to barium oxide, barium hydroxide and

barium hydrate. Future research will focus on the separation of the influence of carbonic acid, if any, in the electrolytic reaction, and identification of the specific salt form under examination. The concentration levels within an individual rain sample of the salt form are expected to be extremely low as indicated by the pH values recorded, and therefore potentially difficult to detect. It is also worth noting that the safe levels of elements such as barium in the environment are also quite low, and on the order of 1-2ppm.

Additional photographs and specifics of the electrochemical testing currently underway will be presented as time and circumstances permit. It is important that the electrodes used be relatively inert with respect to water, such as copper, platinum, silver, or gold. The voltage source currently being used is 24 volts. These preliminary results are presented to expedite additional tests that may soon be conducted by involved citizens. This page will be edited as conditions and information require.

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