20 TIMES

Clifford E Carnicom
September 21 2000

Recent and preliminary pH test data from across the nation indicates that the atmospheric chemistry has been altered by a factor of 20 with respect to hydroxide ion concentration, relative to baseline values established from the years of 1990 -1999, as well as individually with respect to 1999 data. This is a remarkable change in a relatively short period of time, and has major implications for both the chemistry and biology of the nation and the globe.

Human biology is sensitive to pH blood changes as low as 0.1 (approximately 25% change in the H+ concentration); the current data indicates an average change in the pH of rainwaters that are being analyzed

across the country at 1.30. It is important that more citizens become immediately involved in this testing process, as it is relatively inexpensive and simple to accomplish. Significant variations, such as those being currently observed, demonstrate the need for immediate formal investigation into the atmospheric chemistry changes associated with aircraft aerosol operations since the beginning of 1999.

The preliminary data as of this date indicates a 20 times increase in the number of hydroxide ions in the atmosphere, this change apparently occurring primarily within the last year. The following statement from the Nobel Prize winner of Chemistry in 1995 is repeated to emphasize the significance of this topic and the need for public testing and disclosure:

"The single most

important chemical species in clouds and precipitation is the hydrogen ion (H+), whose concentration can be indicated by specifying the solution's acidity, or pH value. You may recall from high school chemistry that the pH scale ranges from 0 to 14, low pH values indicating high acidity (high concentrations of H+) and high pH values indicating high alkalinity (low concentrations of H+)"

from Atmosphere, Climate, and Change by Graedel and Crutzen, Scientific American, 1997.

Please also refer to:

A Case For Testing
pH Test Alert
ph Test Results

Appendix:

The above calculation is based upon the following

definition of pH:

pH = log (1/H+)

Therefore, for two independent pH readings:

pH2 - pH1 = log (1/H2+) - log (1/H1+)

pH2 - pH1 = log ((1/H2+) / (1/H1+))

pH2 - pH1 = log (H1+ / H2+)

or

10^(pH2 - pH1) = H1+ / H2+

and with the average difference in pH being reported as +1.30 as of this date with respect to 1999:

10^1.30 = 20.0

and note that POH is

similarly defined as:

pOH = log (1 / OH-)
leading to similiar
results for the analysis
of hydroxide ions.

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This page will be edited
as circumstances or
conditions require.

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