

ATMOSPHERIC MAGNESIUM DISCLOSED

**Clifford E Carnicom
Jun 10 2001**

The presence of substantial amounts of elemental magnesium within the atmosphere, as procured from rainfall samples, has now been established through recent tests. There is now a need for the extensive collection of rainfall samples by citizens to validate or refute those results which have been obtained locally in Santa Fe, New Mexico.

The method developed to establish this finding is simple to perform, inexpensive and is readily available to all concerned citizens.

The current findings deserve broad distribution due to the simplicity of the method used and the level of access for confirmation or refutation now available to each citizen.

The method of determination of magnesium within atmospheric samples is as follows:

- 1. Collect rainfall samples in a clean container.**

2. Place approximately 2-3 ml (approx. 1/2 tsp.) of a rainfall sample in a clean test tube or container.

3. Add several drops (e.g., 4 to 5) of copper sulfate solution to the rainfall sample. This solution is inexpensive and readily available at aquarium shops.

4. After a few minutes have elapsed, extract a single drop of the rainfall + copper sulfate solution, and place it on a clean microscope slide or a piece of glass.

5. Let the drop evaporate completely at room temperature without disturbance.

6. A crystalline form is expected to develop if sufficient metal concentration exists within the rainfall sample. Each trial conducted thus far on samples obtained Oct 12 2000 and Jun 07 2001 in Santa Fe NM has produced the visible crystalline form.

7. This crystalline form can be identified through a variety of methods, e.g., visual and chemical.

Identification processes thus far confirm the existence of magnesium sulfate as a product of the reaction with copper sulfate. This demonstrates the presence of elemental magnesium within the rainfall sample tested.

Essentially, the magnesium will replace the copper within the solution, and will precipitate into a solid sulfate form that can be crystallized under the conditions described.

8. A ring of copper sulfate crystals,

visually distinctive, is expected to form on the perimeter of the crystal development. Magnesium sulfate crystals will form interior to this ring. 9. An inexpensive hobby microscope (\$15-\$30) at low power (100x) is sufficient to examine the crystal formations.

Numerous control tests of crystalline observations have been conducted, which include:

1. Pure water (reverse osmosis).
2. Reverse osmosis water with copper sulfate added.
3. Copper sulfate solution alone.
4. Rainwater samples alone.
5. Rainwater samples with copper sulfate added.
6. Magnesium sulfate solution alone.
7. Magnesium sulfate with copper sulfate added.

Sufficient variations exist between the control samples and the current findings to justify this presentation. If any alternate identification of the crystal form that develops is made, please provide that finding along with the basis of the decisions.

The reaction under consideration is:



Magnesium sulfate (commonly known as epsom salt) has the following

properties:

1. **Distinctive visual identification under the microscope. Linear, prism or filament crystals are commonly formed.**
2. **Soluble in water.**
3. **Soluble in alcohol.**
4. **Soluble in glycerin.**
5. **Distinctive bitter taste(also known as bitter salt).**
6. **Dissolves in both hydrochloric and sulfuric acid.**
7. **Colorless, transparent or whitish crystal.**

Additional important properties to consider for additional analysis within the aerosol operations include the density, conductivity and ionizability of the element or compound. Magnesium is an extremely light metal, approximately 2/3 the weight of aluminum. Magnesium is extremely conductive, on par with copper and aluminum. Magnesium can be ionized with the energy available within the ultra-violet portion of the spectrum.

These findings provide further evidence of the aerosol operations that have been conducted and remain in progress without the informed consent and knowledge of the citizenry. These findings provide further just cause for the call of accountability and disclosure that remains. All citizens are urged to take an active role to

resolve these issues.

It can be stated that the original motivation for this research involves an attempt to physically identify the barium presence that is strongly evidenced by data that has been made available. The high level of solubility in water of the resulting crystal form immediately dismisses barium sulfate as the candidate of examination for this test that has been developed. Elemental magnesium, combining with the sulfate ion provided by the copper sulfate solution, exists as the viable solution to the problem of identification of this crystal. Sufficient and credible evidence to support the claims of unexpected levels of barium in the atmosphere as an adjunct of the aerosol operations remains in force. Recall that attention has repeatedly been directed toward all elements within Groups I and II of the periodic table.

Past external data that has been made available to me reveals the unusual presence of both magnesium and barium within atmospheric samples. As the data within those reports has not yet been publicly distributed, they have not been given undue attention. Previous data made available to me that indicated the presence of unusual amounts of magnesium in the atmosphere is now elevated in status.

Evidence continues to accumulate that certain metals, i.e., magnesium and barium, as well as certain biological and fibrous components, are established as the core elements of the aerosol operations in progress.

Current testing suggests that magnesium may indeed be the dominant metallic component present. Toxicity levels between varying metals stands as a separate issue. Aggressive testing for these two metal forms as a minimum is now required.

Research on this finding will continue. Any revisions or corrections to this presentation will be made as is appropriate.

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Additional Notes:

Numerous other sulfate forms have been considered as a part of the analysis, and each has been eliminated as a candidate of further evaluation by a variety of methods.

In each of the following cases, one or more conditions preclude further consideration in light of the properties that have been observed. The alternative sulfate forms which have been eliminated from further

consideration within this test form
include:

1. Cadmium sulfate - effloresces, i.e., crumbles upon exposure to the air.
2. Sodium sulfate - (1) visibly different from the crystal obtained (2) yellow solution formed upon addition of copper sulfate to sodium chloride solution. (3) color variation.
3. stannous sulfate - no significant reaction will take place between tin and copper in solution due to the electrochemical potential difference between the elements.
4. barium sulfate - not soluble in water.
5. calcium sulfate - only moderately soluble in water.
6. cobalt sulfate - melting point of 98.6 deg. C.
7. copper sulfate - appearance, blue color.
8. iron sulfate - blue-green color.
9. mercuric sulfate - not soluble in water.
10. magnesium sulfate - not eliminated, satisfies all conditions and properties.
11. manganese sulfate - color, melts at 30 deg. C.
12. lead sulfate - insoluble in alcohol.
13. strontium sulfate - insoluble in alcohol.
14. potassium sulfate - insoluble in alcohol.
15. aluminum sulfate - melting point of 87 - 92.5 deg. C.

The copper sulfate solution used in

The copper sulfate solution used in these tests is manufactured by Aquarium Products, Glen Burnie MD, and has a stated strength of metallic copper at 1.61%.

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